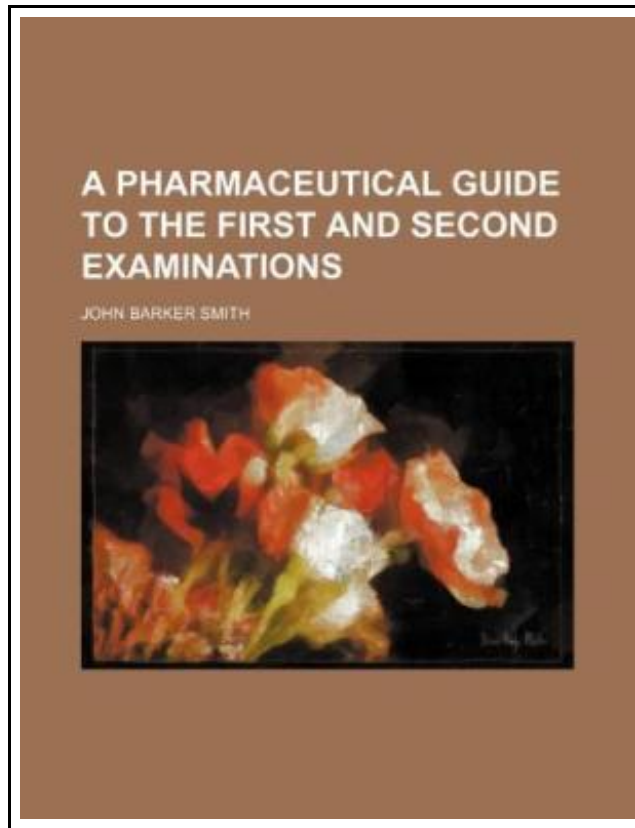


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(Dr. Bethany Lindgren)

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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1870 Excerpt: . to a moderate heat until the nitric oxide be driven off. Magnesium, light; Magnesium, Zinc, Lime.--The light carbonate of the first and the carbonates of the others are heated to a dull red heat in a loosely-covered crucible until there is no effervescence when mixed with acids. Lime is mostly obtained by calcination of limestone on a large scale. 2. Precipitated oxides. Antimony.--A solution of the chloride is partly decomposed by addition of water, and finally by sodium carbonate, the precipitated oxide is then washed with boiling water and dried at a heat not exceeding 212. Iron, magnetic; Iron, peroxide, moist; Iron, hydrated peroxide.--The magnetic is formed on boiling a solution of iron sulphate and persulphate with soda solution; it must be dried at a heat less than 120: the humid peroxide is precipitated with soda solution, but without heat and from a solution of iron persulphate, it is preserved moist in a well-stopped bottle, and is ordered to be recently made; the hydrated is prepared by drying the humid at a temperature not exceeding 212. Silver.--A solution of the nitrate is decomposed by agitation with lime-water, and the resulting oxide dried at a heat not exceeding 212. The black and yellow lotions of mercury are familiar examples of the formation of oxides by addition of a hydrate to metallic salts, the hydrate in these examples being lime-water. 3. Various. Ether; Ether, pure.--Ether is procured by distilling rectified spirit, ethyl hydrate, with sulphuric acid at...

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